

Synthesis, Spectroscopic Characterization, Thermal Analysis and Antibacterial Activity of Ni(II), Cu(II) and Zn(II) Complexes with Schiff bases Derived from β -Diketones

Razieh Ahmadzadeh,¹ Mohammad Azarkish,² and Tahereh Sedaghat^{3*}

¹ Department of Chemistry, Science and Research Branch, Islamic Azad University, Khouzestan, Iran.

² Department of Chemistry, Shoushtar Branch, Islamic Azad University, Shoushtar, Iran.

³ Department of Chemistry, College of Sciences, Shahid Chamran University, Ahvaz, Iran. tsedaghat@scu.ac.ir

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Abstract. Five transition metal complexes, [CuL^a] (1), [NiL^a] (2), [ZnL^a] (3), [CuL^b] (4) and [NiL^b].EtOH (5) have been synthesized from reaction of Ni(II), Cu(II) and Zn(II) acetate salts with two Schiff bases, 3-(2-hydroxy-5-methylphenylamino)-1,3-diphenylprop-2-en-1-one (H₂L^a) and 3-(2-hydroxy-5-methylphenylimino)-1-phenylbuten-1-one (H₂L^b). On the basis of analytical and spectral data, Schiff base is coordinated to metal as tridentate dianionic ligand via phenolic and enolic oxygens and imine nitrogen. Thermal decomposition of the complexes has been studied by thermogravimetry. The *in vitro* antibacterial activity of Schiff bases and their complexes has been evaluated against Gram-positive (*Bacillus subtilis* and *Staphylococcus aureus*) and Gram-negative (*Escherichia coli* and *Pseudomonas aeruginosa*) bacteria and compared with the standard drugs.

Key words: Schiff base, transition metal, thermogravimetry, antibacterial activity, diketones.

Resumen. Se sintetizaron cinco complejos de metales de transición, [CuL^a] (1), [NiL^a] (2), [ZnL^a] (3), [CuL^b] (4) y [NiL^b].EtOH (5), mediante la reacción de iones Ni(II), Cu(II), Zn(II) con dos compuestos tipo base Schiff, 3-(2-hidroxi-5-metilfenilamino)-1,3-difenilprop-2-en-1-ona (H₂L^a) y 3-(2-hidroxi-5-metilfenilimino)-1-fenilbuten-1-ona (H₂L^b). Los resultados de los análisis y los datos espectrales obtenidos revelaron que en la base Schiff actúa como un ligando tridentado dianiónico, coordinando al ion metálico mediante oxígenos fenólico y enólico y mediante nitrógeno del grupo imino. La descomposición térmica de los complejos se estudió por termogravimetría. La actividad antibacteriana de los dos ligandos y de los cinco complejos ha sido evaluada en ensayos *in vitro* utilizando cepas de bacterias Gram-positivas (*Bacillus subtilis*, *Staphylococcus aureus*) y de bacterias Gram-negativas (*Escherichia coli*, *Pseudomonas aeruginosa*) comparando los resultados con los obtenidos para fármacos estándar.

Palabras clave: Base Schiff, metal de transición, termogravimetría, actividad antibacteriana, dicetonas.

Introduction

Schiff bases are among the most widely used ligands and play an important role in metal coordination chemistry due to facile synthesis, strong coordination abilities, remarkable versatility and good biological activities [1]. Many of these compounds exhibit tautomeric rearrangements and receive interest due to several applications in optical recording technology and molecular electronics [2, 3]. The research field dealing with metal complexes of Schiff bases is very broad because of their potential interest for a number of areas including bioinorganic chemistry, catalysis and electrochemistry [4-7]. The development in the field of biological inorganic chemistry has increased the interest in Schiff base complexes. A large number of Schiff bases and their metal complexes have been investigated due to their interesting biological properties such as their ability to reversibly bind oxygen, biological modeling applications and antibacterial, antifungal, anticancer, and herbicidal activities [8-13]. In many cases, when Schiff bases administered as their metal complexes, the biological activity of these complexes is enhanced in comparison to the free ligand [14-17].

Recently we have reported synthesis and antibacterial activities of several Schiff bases derived from β -diketones and their metal complexes [16-20]. These type of Schiff bases contain the N(sp³)-C(sp²)-C(sp²)-C(sp²)=O(sp²) fragment which

is bounded by strong hydrogen bonds (N-H...O or N...H-O). These compounds are of biological interest, play as synthetic intermediates in organic reactions and act as sensor materials [2, 21]. As an extension of our investigation, herein we report synthesis, characterization, thermal behavior and antibacterial activities of Ni(II), Cu(II) and Zn(II) complexes with two Schiff bases, 3-(2-hydroxy-5-methylphenylamino)-1,3-diphenylprop-2-en-1-one (H₂L^a) and 3-(2-hydroxy-5-methylphenylimino)-1-phenylbuten-1-one (H₂L^b) (Fig. 1.).

Results and Discussion

The Schiff bases, 3-(2-hydroxy-5-methylphenylamino)-1,3-diphenylprop-2-en-1-one (H₂L^a) and 3-(2-hydroxy-5-methylphenylimino)-1-phenylbuten-1-one (H₂L^b), have been prepared from reaction of 2-amino-4-methylphenol with dibenzoyl methane or benzoylacetone, respectively. In these reactions only 1:1 condensation was occurred even when an excess of the amine was used. Structural formula for these Schiff bases in three tautomeric forms is given in Fig. 1. According to previous evidences, keto-amine tautomeric form (III) is preferred in solid state [16-19, 22]. The new complexes 1-5 were synthesized by reaction Cu(II), Ni(II) and Zn(II) acetate salts with corresponding Schiff base ligands in ethanol. Acetate anion acts

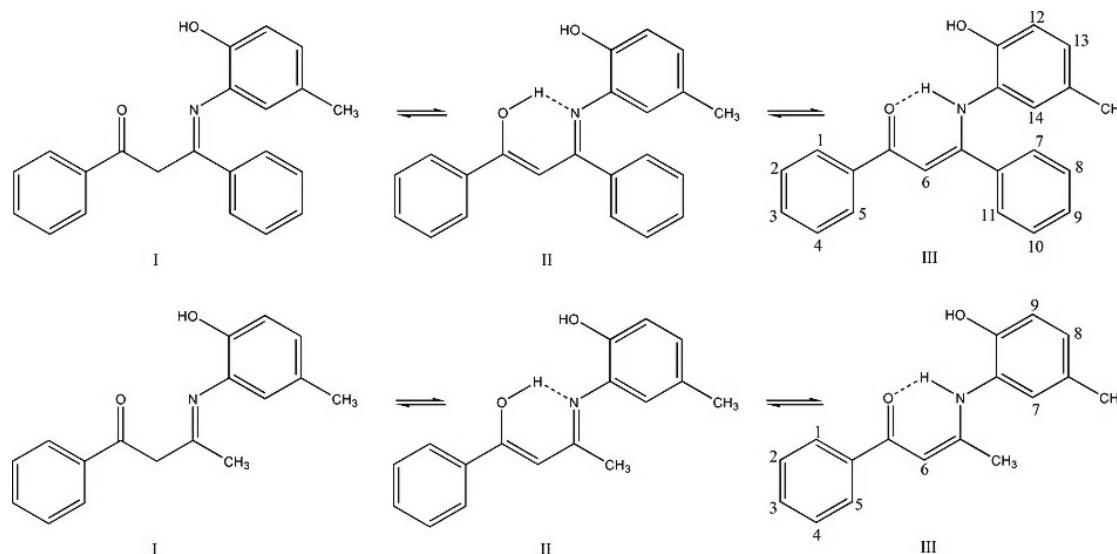


Fig. 1. Tautomeric forms of H_2L^a (above) and H_2L^b (down) with numbering for NMR assignments.

also as a base to force the deprotonation of the ligand. The new complexes were investigated by elemental analysis, spectroscopic methods and thermogravimetric analysis. Our attempts to grow single crystals suitable for X-ray crystallography were abortive.

Spectroscopic studies

Fig. 2 show the IR spectra of ligands and their complexes. In the infrared spectrum of free Schiff bases the stretching vibration of the OH/NH appears at lower frequency and overlap with the $\nu(C-H)$ in the range of $3000-3200\text{ cm}^{-1}$ due to the inter-/intra-molecular hydrogen bonding. This broad band is completely absent in the spectra of all complexes supporting deprotonation of the ligand during coordination. In the infrared spectra of all complexes no band is observed in the region $3300-3500\text{ cm}^{-1}$ attributable to the stretching vibration of hydrated or coordinated water. In the spectrum of free ligands a band observed in 1615 cm^{-1} assigned as a perturbed carbonyl stretching with the frequency lowering from a free carbonyl ascribed to conjugation and hydrogen bonding in keto-amine form (III). In the IR spectra of complexes, this band shifts to lower frequency providing evidence of participation of oxygen in bonding with metal and so weakening of C=O bond. A band at $1583-1590\text{ cm}^{-1}$ in the spectra of complexes assigned to C=C/C=N bond indicates the anionic ligand is coordinated to metal in the tautomeric form II. The appearance of new bands in the IR spectra of the complexes in the region $453-585\text{ cm}^{-1}$, which may be assigned to $\nu(M-N)$ and $\nu(M-O)$, supports the bonding of nitrogen and oxygen to the metal ion [23-25].

In the electronic spectra of complexes, d-d bands were not observed because of low solubility of compounds in DMSO and small molar absorption coefficient of these forbidden transitions. Therefore all appeared bands in electronic spectra result from the overlap of $\pi \rightarrow \pi^*$ transitions mainly localized within the imine chromophore and the ligand to metal charge-transfer

(LMCT) transition from the lone pairs of the oxygen to the M(II) ions [26]. 1H NMR spectrum for Zn(II) complex was not recorded because of low solubility of this complex in common NMR solvents.

Thermogravimetric analyses

The thermogravimetric analyses of complexes have been studied from ambient temperature up to $1000\text{ }^\circ\text{C}$ under a N_2 atmosphere. The thermogravimetric and derivative thermogravimetric (TG and DTG) analysis curves for **1-5** are represented in Fig. 3. Thermal behaviors of complexes have been summarized in Table 1. The results show the presence of a ligand molecule per metal ion and confirm the formulae suggested from the analytical data. The absence of weight loss up to $200\text{ }^\circ\text{C}$ indicates that there is no hydrated water molecule in the crystalline solid in **1-4** complexes, while complex **5** losses one mol EtOH in this range. Complex **1** experience a weight loss between $202-620\text{ }^\circ\text{C}$ which can be interpreted as loss of the greatest part of the ligand. Complex **2** shows a drastic mass loss in one step within temperature range $299-855\text{ }^\circ\text{C}$ which attributes to the decomposition of the ligand and the removal of the $C_{17}H_{17}NO$. Complex **3** also shows one-step decomposition between $235-860\text{ }^\circ\text{C}$ corresponds to similar mass loss. The thermal degradation of complex **4** corresponds to the loss of $C_{14}H_{15}NO$ in one step. In the first thermal degradation step of **5**, ethanol is eliminated and the second step ($325-540\text{ }^\circ\text{C}$) evolves $C_{12}H_{15}NO$. In all cases the residues are carbon and metal oxide. In general, complexes **1-5** exhibit good thermal stability and the Ni(II) complexes are more stable than the others.

Biological studies

The *in vitro* antibacterial activities of Schiff bases and their complexes were studied along with two standard antibacterial drugs, *viz.* Nalidixic acid and Vancomycin. The microorgan-

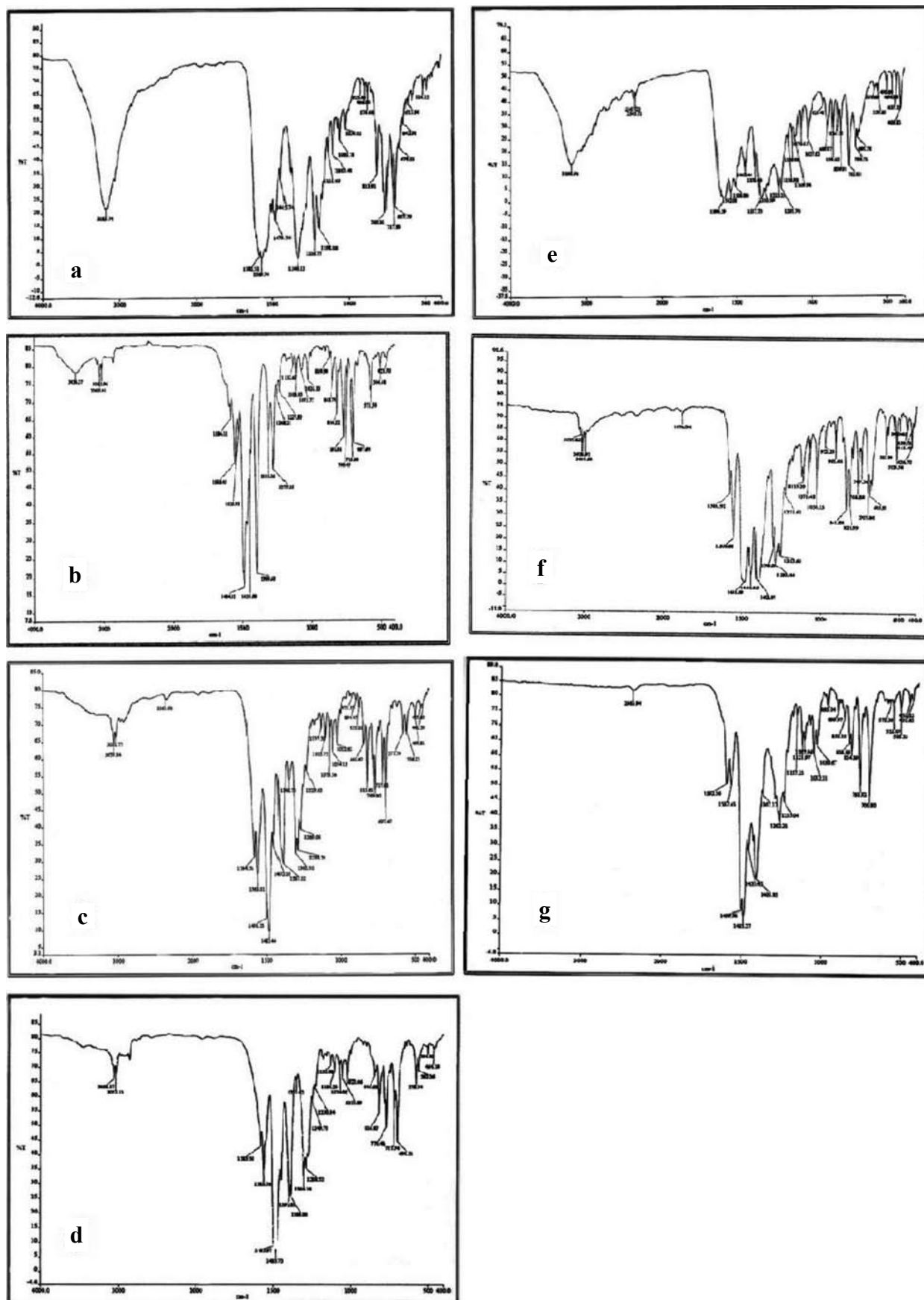


Fig. 2. IR spectra of (a) H₂L^a, (b) 1, (c) 2, (d) 3, (e) H₂L^b, (f) 4, (g) 5.

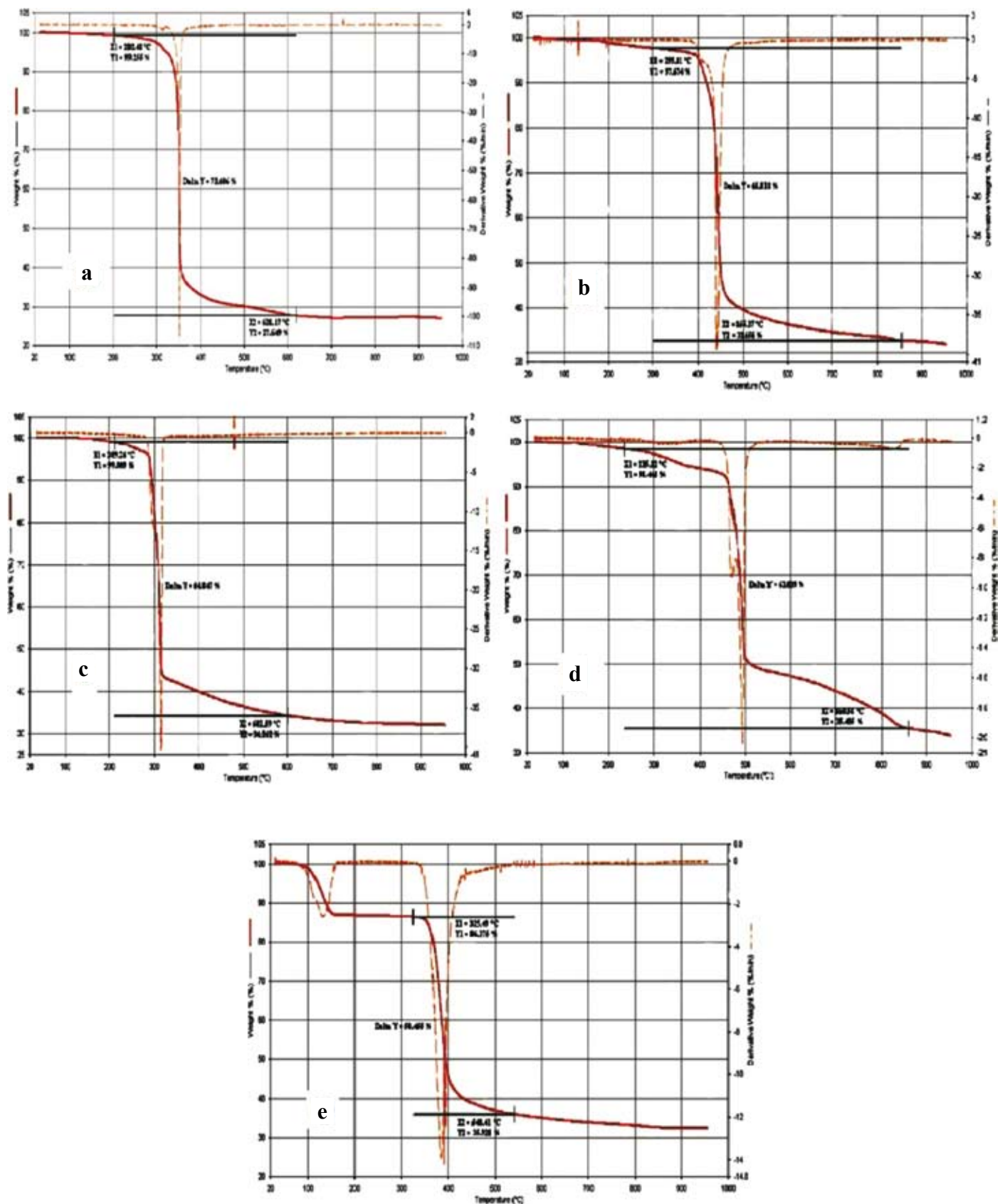


Fig. 3. TG and DTG curve of synthesized complexes: (a) 1, (b) 2, (c) 3, (d) 4 and (e) 5.

isms used in this work include *Bacillus subtilis* and *Staphylococcus aureus* (as Gram-positive bacteria) and *Escherichia coli* and *Pseudomonas aeruginosa* (as Gram-negative bacteria). The results are presented in Table 2. The antibacterial activity of compounds is due to either bactericide effects (killing the bacteria) or bacteriostatic effects (inhibiting multiplication of bacteria by blocking their active sites). It may be postulated that antibacterial compounds deactivate various cellular enzymes, which play a vital role in various metabolic pathways of these organisms. It has also been proposed that the ultimate action of the antibacterial agent is the denaturation of one or more proteins of the cell, which, as a result, impairs normal cellular processes [27]. However, it is apparent that permeability across

the bacterial cell wall is necessary for the effectiveness of the biocide compounds; therefore any factor which facilitates microorganism membrane crossing may enhance the activity. Comparing the biological activity of the Schiff bases, complexes and standard drugs, indicate that all complexes exhibit more inhibitory effects than the parent ligands against bacterial strains. This enhanced antibacterial activity may be due to electron delocalization over the whole chelate ring upon complexation. Such chelation increases the lipophilicity and enhances the permeation through the lipid layer of the cell membrane (chelation theory) [28-32]. It is apparent that complexes were more toxic towards Gram-positive than Gram-negative bacteria strains. The reason is the difference in the structures of the cell walls. Lipopolysaccharides form an outer lipid membrane and contribute to the complex antigenic specificity of Gram-negative cells [33].

Table 1. Thermal decomposition data for complexes 1-3.

| Complex | Temp. range (°C) | DTG peak temp. (°C) | Weight loss(%) obs.(calcd.) | Evolved product |
|---------|------------------|---------------------|-----------------------------|------------------------------------|
| 1 | 202-620 | 352 | 71.60 (70.40) | C ₁₉ H ₁₇ NO |
| 2 | 299-855 | 441 | 65.02 (65.01) | C ₁₇ H ₁₇ NO |
| 3 | 235-860 | 493 | 63.04 (63.90) | C ₁₇ H ₁₇ NO |
| 4 | 209-602 | 315 | 64.84 (64.77) | C ₁₄ H ₁₇ NO |
| 5 | 95-325 | 130 | 13.62 (12.44) | C ₂ H ₆ O |
| | 325-540 | 391 | 50.45 (51.12) | C ₁₂ H ₁₅ NO |

Conclusion

On the basis of above analytical and spectral data, in all complexes the Schiff base is deprotonated and coordinated to metal ion as tridentate dianionic ligand via imine nitrogen and phenolic and enolic oxygens. Thermal and elemental analysis data show no water molecules in the formula of complexes. It is suggested that the coordination sphere of complex is completed by

Table 2. Antibacterial activity data of ligands and their complexes.

| Compound | Conc. (µg/ml) | Inhibition zone (mm) | | | |
|-------------------------------|---------------|----------------------|----------------------|------------------|--------------------|
| | | <i>E. coli</i> | <i>P. aeruginosa</i> | <i>S. aureus</i> | <i>B. subtilis</i> |
| H ₂ L ^a | 100 | n.a. | n.a. | n.a. | n.a. |
| | 200 | n.a. | n.a. | n.a. | n.a. |
| | 400 | n.a. | n.a. | n.a. | n.a. |
| H ₂ L ^b | 100 | 9 | n.a. | n.a. | n.a. |
| | 200 | 10 | n.a. | n.a. | n.a. |
| | 400 | 13 | n.a. | n.a. | n.a. |
| 1 | 100 | 12 | n.a. | 14 | 16 |
| | 200 | 13 | n.a. | 17 | 17 |
| | 400 | 16 | n.a. | 19 | 20 |
| 2 | 100 | 9 | n.a. | 9 | 13 |
| | 200 | 15 | 11 | 13 | 17 |
| | 400 | 17 | 12 | 18 | 19 |
| 3 | 100 | 11 | n.a. | 17 | 15 |
| | 200 | 14 | 13 | 18 | 17 |
| | 400 | 16 | 14 | 20 | 19 |
| 4 | 100 | 11 | n.a. | 18 | 11 |
| | 200 | 12 | 11 | 20 | 16 |
| | 400 | 15 | 13 | 24 | 18 |
| 5 | 100 | n.a. | n.a. | 19 | 18 |
| | 200 | n.a. | n.a. | 21 | 20 |
| | 400 | 10 | 10 | 22 | 23 |
| Vancomycin | | 22 | 8 | 16 | 15 |
| Nalidixic acid | | 28 | 10 | 11 | 17 |

n.a. = no activity

phenolic oxygen atom of another molecule generating a dimer. Similar dimeric structures have been reported earlier for Schiff base complexes [19, 28, 34-38]. According to previous report [20], the interaction occurs with the oxygen atom forming a part of the five-membered heterocycle that can be attributed to the increased "s" character at the five-membered ring [37]. Therefore, we suggest a dimeric structure for synthesized complexes with a distorted square planar geometry around metal ions (Fig. 4). On the basis of thermal analysis data, the complexes have been found to be thermally stable. All complexes exhibited good activities against gram positive bacteria and have a potential to be used as drugs.

Experimental

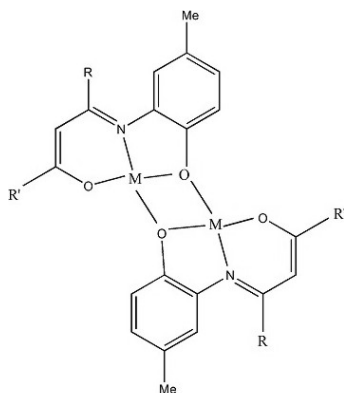
Materials and methods

All starting materials were purchased from Merck Company and used as received. All solvents were of reagent grade and used without further purification. IR spectra were obtained using a FT BOMEM MB102 spectrophotometer. The ^1H NMR spectra were recorded in DMSO- d_6 with a Bruker 400 MHz Avance Ultrashield spectrometer. Thermogravimetric analyses (TGA) were carried out using a Perkin-Elmer Diamond thermal analysis. The heating rates were controlled at 5°C min^{-1} under a nitrogen atmosphere with a 150 ml/min flow rate and the weight loss was measured from ambient temperature up to 1000°C .

Synthesis of Schiff bases

3-(2-hydroxy-5-methylphenylamino)-1,3-diphenylprop-2-en-1-one (H₂L^a)

H_2L^a was prepared from reaction of 2-amino-4-methylphenol and dibenzoylmethane as the method reported earlier [19], herein some ^1H NMR data have been revised. λ_{max} (nm, DMSO):



- (1) M = Cu, R = Ph, R' = Ph
 (2) M = Ni, R = Ph, R' = Ph
 (3) M = Zn, R = Ph, R' = Ph
 (4) M = Cu, R = Me, R' = Ph
 (5) M = Ni, R = Me, R' = Ph

Fig. 4. Suggested structure for synthesized complexes.

400, 380, 360; IR (KBr, cm^{-1}): 1589, $\nu(\text{C}=\text{C})$; 1615, $\nu(\text{C}=\text{O})$, 3000-3150 br, $\nu(\text{C}-\text{H})$, $\nu(\text{O}-\text{H})/\nu(\text{N}-\text{H})$. ^1H NMR (DMSO- d_6): δ 1.86 (s, 3H, CH_3), 6.08 (d, $^4J = 1.3$ Hz, 1H, H_{14}), 6.11 (s, 1H, H_6), 6.65 (dd, $^3J = 8.0$, $^4J = 1.5$ Hz, 1H, H_{13}), 6.76 (d, $^3J = 8.1$ Hz, 1H, H_{12}), 7.39-7.55 (m, 8H, $\text{H}_{2,3,4,7,8,9,10,11}$), 7.99 (d, $^3J = 7.2$ Hz, 2H, $\text{H}_{1,5}$), 9.79 (s, 1H, $\text{OH}_{\text{phenolic}}$), 12.61 (s, 1H, $\text{NH/OH}_{\text{enolic}}$).

3-(2-hydroxy-5-methylphenylimino)-1-phenylbuten-1-one (H₂L^b)

An ethanolic solution (20 ml) of 2-amino-4-methylphenol (1.232 g, 10 mmol) was added to a solution of benzoylacetone (1.622 g, 10 mmol) in ethanol (20 ml). This solution was refluxed for 8 h. The Schiff base precipitates after standing 24 h at rt. The product was filtered off and washed with ethanol (2×5 mL). Yield: 1.944 g (73%); m.p. $148-150^\circ\text{C}$; λ_{max} (nm, DMSO): 368; IR (KBr, cm^{-1}): 1589, $\nu(\text{C}=\text{C})$; 1615, $\nu(\text{C}=\text{O})$, 3000-3150 br, $\nu(\text{C}-\text{H})$, $\nu(\text{O}-\text{H})/\nu(\text{N}-\text{H})$. ^1H NMR (DMSO- d_6): δ 2.15 (s, 3H, CH_3), 2.23 (s, 3H, CH_3), 6.03 (s, 1H, H_6), 6.87 (m, 2H, $\text{H}_{8,9}$), 7.08 (s, 1H, H_7), 7.45 (m, 2H, $\text{H}_{2,3,4}$), 7.92 (d, $^3J = 6.4$ Hz, 2H, $\text{H}_{1,5}$), 9.80 (s, 1H, $\text{OH}_{\text{phenolic}}$), 12.82 (s, 1H, $\text{NH/OH}_{\text{enolic}}$).

General procedure for synthesis of complexes

The Schiff base (0.5 mmol) in EtOH (5 ml) was added to an equimolar amount of $\text{Ni}(\text{OAc})_2 \cdot 4\text{H}_2\text{O}$, $\text{Cu}(\text{OAc})_2 \cdot \text{H}_2\text{O}$ or $\text{Zn}(\text{OAc})_2 \cdot 2\text{H}_2\text{O}$ in EtOH (5 ml). The solution was refluxed about 4 h for **1**, **2** and **3**, whereas 8 h for **4**. Complex **5** was precipitate after 3 h stirring in rt. The products were filtered, washed with MeOH and dried in vacuum on CaCl_2 . Our attempts to synthesize $[\text{ZnL}^b]$ were unsuccessful.

$[\text{CuL}^a]$ (**1**): Yield: 0.249 g (64%); λ_{max} (nm, DMSO): 317, 436; IR (KBr, cm^{-1}): $\nu(\text{C}=\text{N}) / (\text{C}=\text{C})$, 1584; $\nu(\text{C}-\text{O})$, 1553; $\nu(\text{Cu}-\text{N})$, 571; $\nu(\text{Cu}-\text{O})$, 455; *Anal.* Calcd. for $\text{C}_{22}\text{H}_{17}\text{NO}_2\text{Cu}$: C, 67.53; H, 4.35; N, 3.56%. Found: C, 67.97; H, 4.13; N, 3.94%.

$[\text{NiL}^a]$ (**2**): Yield: 0.236 g (61%); λ_{max} (nm, DMSO): 458, 300; IR (KBr, cm^{-1}): $\nu(\text{C}=\text{N})/(\text{C}=\text{C})$, 1584; $\nu(\text{C}-\text{O})$, 1563; $\nu(\text{Ni}-\text{N})$, 556; $\nu(\text{Ni}-\text{O})$, 455; *Anal.* Calcd. for $\text{C}_{22}\text{H}_{17}\text{NO}_2\text{Ni}$: C, 68.39; H, 4.40; N, 3.62%. Found: C, 68.26; H, 4.40; N, 3.52%.

$[\text{ZnL}^a]$ (**3**): Yield: 0.231 g (62%); λ_{max} (nm, DMSO): 351, 452; IR (KBr, cm^{-1}): $\nu(\text{C}=\text{N})/(\text{C}=\text{C})$, 1583; $\nu(\text{C}-\text{O})$, 1563; $\nu(\text{Zn}-\text{N})$, 562; $\nu(\text{Zn}-\text{O})$, 464; *Anal.* Calcd. for $\text{C}_{22}\text{H}_{17}\text{NO}_2\text{Zn}$: C, 67.21; H, 4.32; N, 3.56%. Found: C, 66.79; H, 4.37; N, 3.51%.

$[\text{CuL}^b]$ (**4**): Yield: 0.280 g (86%); λ_{max} (nm, DMSO): 435, 360; IR (KBr, cm^{-1}): $\nu(\text{C}=\text{N})/(\text{C}=\text{C})$, 1583; $\nu(\text{C}-\text{O})$, 1556; $\nu(\text{Cu}-\text{N})$, 585; $\nu(\text{Cu}-\text{O})$, 459; *Anal.* Calcd. for $\text{C}_{17}\text{H}_{15}\text{NO}_2\text{Cu}$: C, 62.03; H, 4.56; N, 4.26%. Found: C, 62.15; H, 4.48; N, 4.3%.

$[\text{NiL}^b]$ (**5**): Yield: 0.243 g (75%); λ_{max} (nm, DMSO): 439, 310; IR (KBr, cm^{-1}): $\nu(\text{C}=\text{N})/(\text{C}=\text{C})$, 1592; $\nu(\text{C}-\text{O})$, 1567; $\nu(\text{Ni}-\text{N})$, 572; $\nu(\text{Ni}-\text{O})$, 453; *Anal.* Calcd. for $\text{C}_{19}\text{H}_{21}\text{NO}_3\text{Ni}$: C, 61.67; H, 5.68; N, 3.78%. Found: C, 61.45; H, 5.70; N, 3.91%.

Antibacterial tests

The *in vitro* antibacterial activity of ligands and their corresponding complexes was investigated against the standard

strains of two Gram-positive (*Bacillus subtilis* ATCC 12711 and *Staphylococcus aureus* ATCC 6538) and two Gram-negative (*Escherichia coli* ATCC 35218 and *Pseudomonas aeruginosa* ATCC 27853) bacteria. In order to compare the results, Nalidixic acid (30 mg/disc) and Vancomycin (30 mg/disc) were used as standard antibacterial drugs. Determination of the antibacterial activity was carried out by paper-disc diffusion method. The compounds were dissolved in DMSO at 100, 200 and 400 µg/ml concentration. Muller Hinton broth was used for preparing basal media for the bioassay of the organisms. A lawn culture from 0.5 MacFarland suspension of each strain was prepared on Muller Hinton agar. Blank paper discs (6.4 mm diameter) were saturated with a solution of test compounds and placed on the surface of the agar plates. On one paper disc only DMSO was poured as a control. The plates were incubated at 37 °C for 24 h. The inhibition zone diameters around each disc were measured in mm.

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